*New South Wales Chemical Education Group*

# NSW Schools Titration Competition

**Team Score:**

| **Centre: SAMPLE RESULTS SHEET** | **School:** |
|---|---|

*In the table below, record at least two and up to three sets of titration data for each acid.*
*Do not include non-concordant data which was rejected.*
*You must not average your NaOH titres between team members.*
*Calculate the concentration of your acetic acid solution to four significant figures.*

*Invitation to the National Competition is at the discretion of the organising committee.*

| **Solution A** was analysed by: | | **Solution B** was analysed by: | | **Solution C** was analysed by: | |
|---|---|---|---|---|---|
| Titres of NaOH against HCl | Titres of NaOH against Acetic Acid | Titres of NaOH against HCl | Titres of NaOH against Acetic Acid | Titres of NaOH against HCl | Titres of NaOH against Acetic Acid |
| 1. mL | 1. mL | 1. mL | 1. mL | 1. mL | 1. mL |
| 2. mL | 2. mL | 2. mL | 2. mL | 2. mL | 2. mL |
| 3. mL | 3. mL | 3. mL | 3. mL | 3. mL | 3. mL |
| The concentration of **Solution A** was: mol $L^{-1}$ | | The concentration of **Solution B** was: mol $L^{-1}$ | | The concentration of **Solution C** was: mol $L^{-1}$ | |